

Preparation and Characterization of CuS-Graphene Composite

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Abstract. A composite of CuS–graphene was prepared via a mild experimental condition, using copper nitrate, EDTA, glucose, carbon disulfide, ethylenediamine and graphene oxide as the raw materials. The composite was characterized by X-ray power diffraction, X-ray photoelectron spectroscopy, transmission electron microscopy and scanning electron microscope. In this process, graphene oxide was reduced to graphene with the formation of CuS.

Introduction

During the last years, materials scientists have been showed enough interest and made great efforts to develop and apply the transition-metal chalcogenides. Transition-metal chalcogenides can be widely used in sensors, catalyst, lithium battery; optical devices, solar cell, and so on ^[1-3]. Copper sulfides, an important member of the transition-metal chalcogenides, is one of the p-type compound semiconductors with narrow band gap. They were potentially important material in optical filter, superionic material, cathode material, catalyst, solarcontroller, etc ^[4, 5]. Copper sulfides with various morphologies have been successfully fabricated involving nanoparticles, nanorods nanowires, nanotrees, nanoplates, hollows pheres, nanoflakes ^[6, 7].

Graphene, as its many excellent properties ^[8, 9], have attracted significant attention. It has been employed to as electrode material, sensors, nanoelectronics, energy storage devices, and capacitors ^[10]. Meanwhile, as a new carbon material, graphene can also be used to prepare the composite material. There were a lot of graphene-based composite materials. Our group has been achieved Ni₃S₂/Graphene, CdS/Graphene, ZnS-Graphene, and Ag₂S-Graphene composite ^[11-14].

Herein, we demonstrate a facile fabrication of graphene–copper sulfides composite via a mild reaction conditions. Throughout the experiment, graphene oxide (GO) was reduced as a support material. Simultaneously, copper sulfides were attached onto the graphene sheets, and prevented the aggregation of the graphene sheets.

Experimental section

Preparation of CuS–Graphene composite

GO was prepared from natural graphite powder according to the method reported by Hummers and Offeman ^[15]. In a typical preparation about CuS–Graphene composite (CuS-G) ^[16]: 20 mg of GO was dispersed into 200 mL of distilled water with sonication for 45 min. Then, 250 mg of Cu (NO₃)₂·3H₂O and 300 mg of EDTA were added with stirring for 3 h at room temperature. Concomitantly, 0.1 M glucose solution was added and stirred for 30 min. afterwards, 1 M NaOH was added and stirred for 10 min. After that, to the above system, the solution was heated to 90 °C in a water bath and kept at this temperature for 1 h. Finally, the solution A was added and stirred for a period of time. The precipitate was collected by centrifugation, and then it was washed with distilled water and absolute ethanol several times and dried in a vacuum oven at 60 °C for 12 h. Solution A ^[13] consisted of Carbon disulfide (CS₂) and ethylenediamine aqueous solution. This solution, in a beaker, was stirred at room temperature till there was no CS₂ droplet found at the

bottom of the container. For comparison, the same method was used to synthesize CuS without GO.

Characterization

The compositions and structures of the samples were characterized by powder X-ray diffraction (XRD) using a Bruker D8 Advance diffractometer with Cu K α radiation ($\lambda=0.15418$ nm). The aerogel microstructure was imaged using JEOL JEM-2100 transmission electron microscopy (TEM). Scanning electron microscope (SEM, HITACHIS-4800) was used to analyze the morphology of the samples. Fourier transform infrared (FTIR) spectra of KBr powder pressed pellets were recorded on a Bruker VECTOR 22 spectrometer in the frequency range of 4000–450 cm^{-1} . X-ray photoelectron spectra (XPS) were carried out on a RBD upgraded PHI-5000C ESCA system (Perkin Elmer) with Mg K α radiation ($h\nu=1253.6$ eV).

Results and Discussion

The as-prepared GO, graphene, CuS and CuS-G product were characterized by XRD. It is seen from fig. 1 that all the patterns of these two powders can be indexed accurately on the standard diffraction data of hexagonal CuS phase (JCPDS No.02-0820, space group: P63/mmc). The 2θ values of the peaks are 27.5, 29.4, 31.8, and 48.1, assigning to (101), (102), (103), and (110) reflections of hexagonal CuS, respectively. However, there is no peak of GO (001) or graphene (002) in composite. According to research reported in our previous, there are several possible causes for this phenomenon. This is because GO was reduced to graphene and the exfoliated graphene show no peak of (002). Moreover, the crystal growth of CuS between the interlayers can destroy the regular layer stacking of GO, resulting the weakening or even disappearance of the GO (001) diffraction peak^[17, 18].

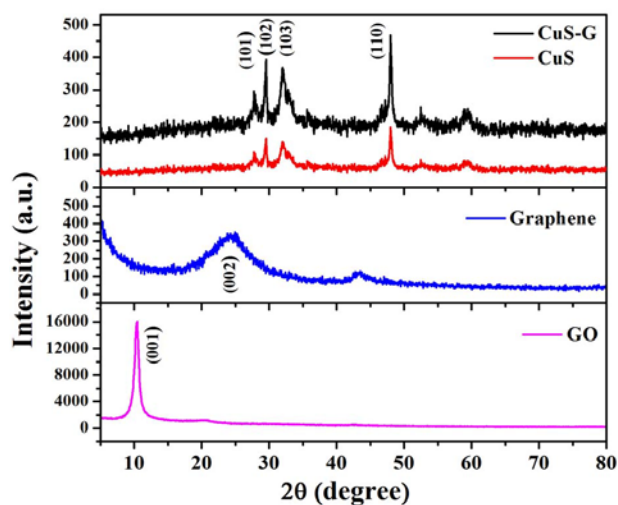


Figure 1 XRD diffraction patterns of GO, graphene, CuS and CuS-G.

The detailed structure of CuS and CuS-G was shown by SEM and TEM in fig. 2. As can be seen from the images, cubic shape was obtained while solution A was added. With the extending of the reaction time, the mixtures of cubic and hollow cube copper sulfides are appears. When the reaction time of over 24 h, the copper sulfide exhibit amorphous structure. In fig. 2b, it can be clearly seen that a great number of cube-shaped products were attached onto the carbon sheets. Simultaneously, some of cubic products were inserted into sheets, but not connected to the surface of sheets. According to the report,^[19, 20] when graphite oxide to grapheme by the chemical conversion, it is prone to holes and defects on the carbon grid. Therefore, the crystal may be grown around these holes and defects, and they tend to be inserted into the sheets^[21]. Base on previous studies about the formation mechanism,^[16, 22] we surmised the possible formation mechanism of CuS-G. Firstly, Cu²⁺ and EDTA can form Cu²⁺-EDTA; then the formed Cu²⁺-EDTA can transfer to the GO sheets via the intercalation or adsorption. In addition, the glucose can be used as the reducing agent. In the reaction process, Cu²⁺ can be reduced to Cu⁺ in the presence of glucose, forming Cu₂O crystals. At

the same time, H_2S can be generated by CS_2 and ethylenediamine in the solution A. After reacting with H_2S , the Cu_2O is converted into CuS crystals. Moreover, throughout the reaction, GO is reduced to graphene.

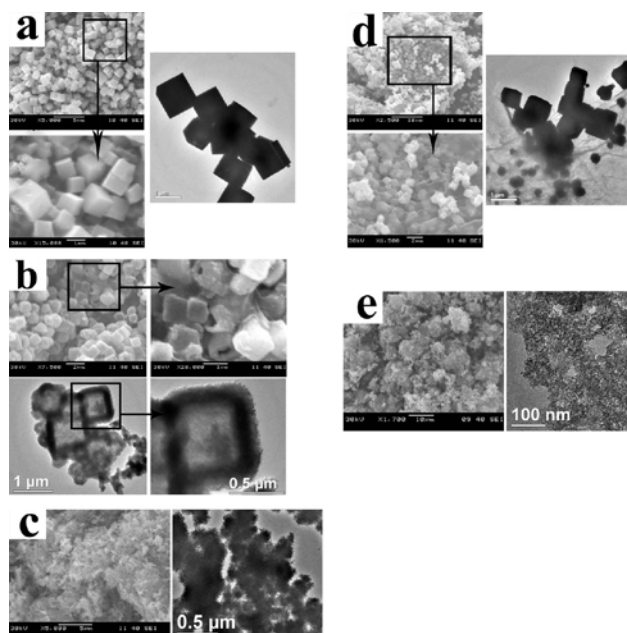


Figure 2 Typical SEM and TEM images of (a) Cu_2O , (b) CuS (after 1h the solution A was added), (c) CuS 5h (after 5h the solution A was added), (d) Cu_2O -G, (e) CuS -G (after 1h the solution A was added)

FTIR is employed to better observe the chemical structure of GO, CuS -G composite. In fig. 3, the characteristic peaks of GO are the absorption bands corresponding to epoxy C–O stretching vibrations at 1042 cm^{-1} , O–H deformation vibrations of tertiary C–OH at 1390 cm^{-1} , carboxy O–H deformation vibrations at 1622 cm^{-1} , and carboxy C=O stretching vibrations at 1729 cm^{-1} . However, the peaks for the oxygen-containing functional groups after the reaction are reduced significantly. Moreover, for CuS -G, the adsorption peak around 1566 cm^{-1} may be attributed to the unoxidized carbon backbone stretching vibrations of graphene^[17, 23]. Simultaneously, the absorption peak at the 940 cm^{-1} associated with Cu–S vibration. The band at 1081 cm^{-1} corresponding to C–H bending frequencies. In addition, the composite can be further testified by XPS data (fig. 4). In XPS spectra, the binding energy at 937.3 eV and 162.1 eV is assigned to Cu $2p_{3/2}$ and S $2p_{1/2}$, respectively.

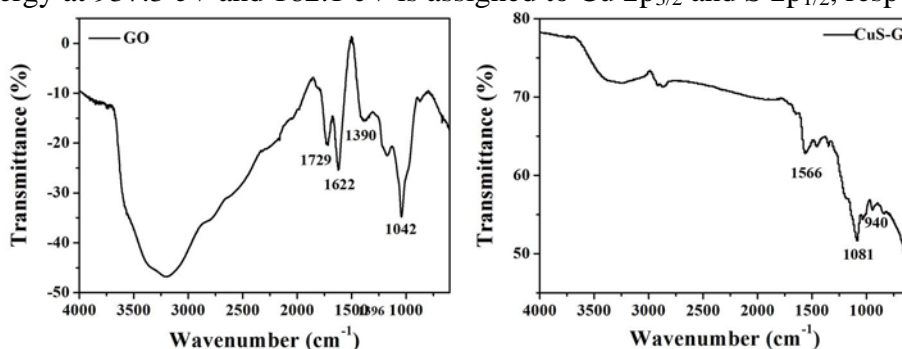


Figure 3 FTIR spectra of GO and CuS -G

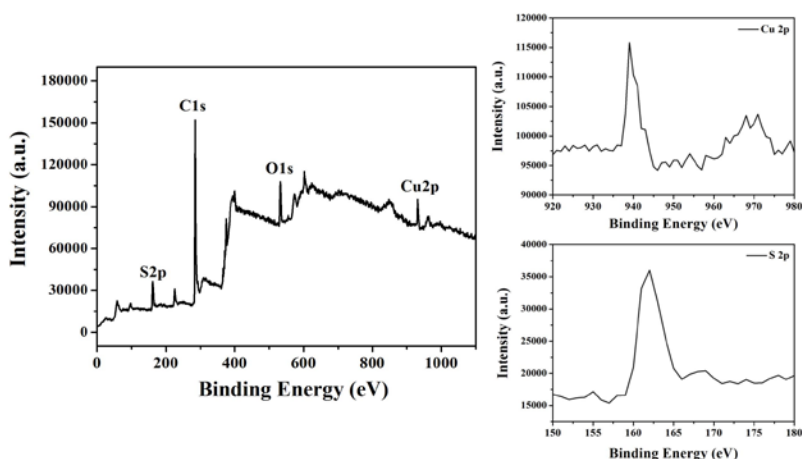


Figure 4 XPS wide-scan spectrum of CuS–G composite and high-resolution XPS spectra of Cu 2p, S 2p

Conclusions

In conclusion, CuS were deposited onto graphene sheets by a mild experimental condition. Experimental results indicate that graphene oxide was reduced to graphene with the formation of CuS.

Acknowledgements

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