

Research Progress of Fluorescent Molecular Probe for Cyanide Ion

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ABSTRACT: Cyanide ion, a kind of extremely toxic chemical, is used as chemical warfare agents, as well as civilian chemical raw stuffs. Many detection methods for cyanides ion have been researched. This paper reviewed on some research progress of fluorescent molecular probe for cyanide ion in recent 5 years in worldwide, as a conclusion, the authors assessed the applicability of fluorescent probes with different identifying mechanisms for cyanides in water.

KEYWORD: Fluorescent molecular probe; cyanide ion; research progress

1 INTRODUCTION

Cyanides are extremely toxic substances, with micro lethal dosage. Except for taking accidentally, the vapor and dust of cyanides can access bodies by respiratory and digestive tract. Cyanides even combine with ferric ion of cytochrome oxidase in vivo by permeating through skins, which lead to the cells losing the ability of adsorbing and delivering oxygen, accordingly oxygen lacked in bodies, eventually causes mammals poisoning and death[1]. However, cyanides are omnipresent in human's daily life, for example, in gold mining, electroplating and metallurgy industry, moreover, cyanides utilized as raw materials of synthetic resin, medicines, pesticides[2]. Cyanides are also the precursors of chemical warfare agent, and they were ever utilized as chemical warfare agent during WWI, at the end of WWII, Japan abandoned a large number of chemical weapons contained cyanides in China[3], which have given rise to severe pollution of soil and water source in many areas. Thus development of trace amount detection method is vital in environment monitoring and medical research.

So far, there have been many reported detection and analysis methods for cyanide ion (CN^-) in literatures, including titrimetric analysis[4], electrochemical process[5], ion chromatography[6], and biosensor[7]. Nevertheless, those methods utilized to identify CN^- are not only time-consuming and operational complex, but also needing high precision instruments. While fluorescent molecular probe method has some advantages of simple operation, good selectivity,

high sensitivity, as well as realizing on-line analysis, immediate analysis and cells imaging even without expensive instruments. In 2009, Huang Zhiping[8] reviews on fluorescent detection of CN^- , Mu Xin[9] reviews on the application of triarylglyhydroboron detection for fluorinon and CN^- . This paper focuses on some worldwide research progress of fluorescent molecular probe for CN^- in recent 5 years, furthermore, assesses the applicability of different mechanisms of fluorescent molecular probe for CN^- in water.

2 RESEARCH PROGRESS OF DETECTION FOR CYANIDE ION

According to identification mechanisms, the acceptor molecules of fluorescent molecular probe for cyanide ion can be classified into 3 types, i.e. acceptor molecules[10] of CN^- being bonded by H-bond or electrostatic interaction; acceptor molecules[11,12] contain metal ion or lewisite acid; reaction type acceptor molecules (so called chemometry recognizer)[13,14]. They are accounted for as follows.

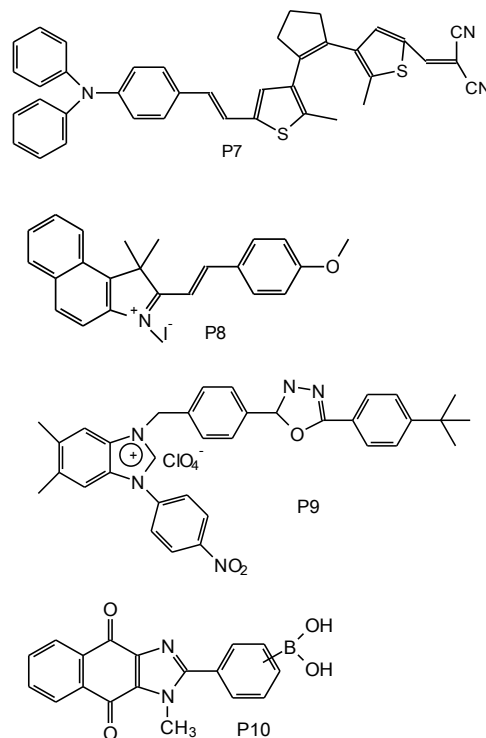
2.1 Fluorescent molecular probe based on H-bond interaction

H-bond interaction is a common used method to identify anion. Compared to other anion, cyanide ion is not a stronger acceptor of H-bond, but it is reported that CN^- can be identified by H-bond. Odago[15] illustrates colorimetric CN^- probe P1

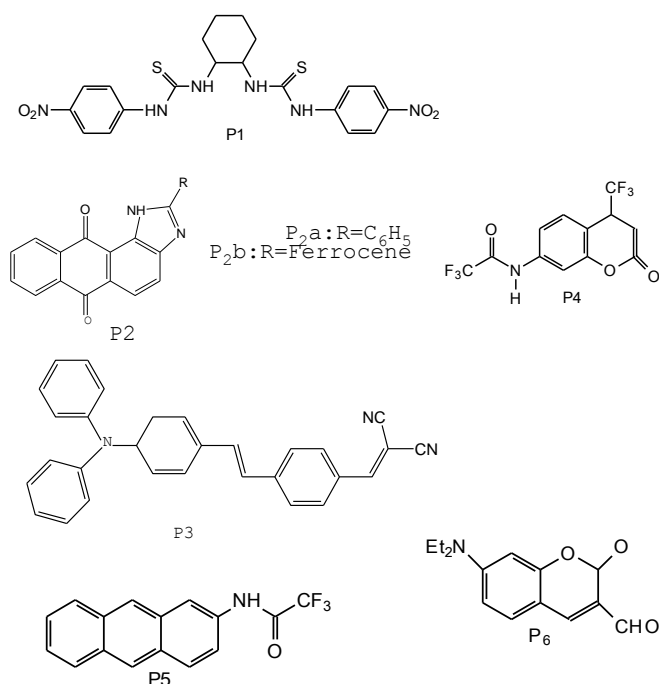
which contains 2 thiocarbamide groups (scheme 1). In DMSO, P1 is pale yellow solution, after CN^- is added, the solution color turns into red, thus CN^- is identified by visual sensing detection. Then polar solvent methanol is added into the solution, solution color recovers from red to yellow. By ^1H NMR, it is verified that the de-proton process of thiocarbamide group is indeed occurred. The mechanism can be explained as: after CN^- is added, the de-proton process of thiocarbamide accelerate the intramolecular charges transferring (ICT). Sukdeb[16] reports probe P2 that contains imidazole group. In HEPES: CH_3CN =1:1 solution, both P2a and P2b can identify CN^- , the color of P2a turns from pale yellow to red, while P2b turns from red to dark blue. Moreover, P2a also can identify CN^- by fluorescent ratio. The minimum detection limit of P2a is 0.06mg/L, while P2b is 0.078mg/L, which are approximate to the WHO's drinking water standard what of 0.05mg/L[17].

2.2 Fluorescent molecular probe based on metal coordination compounds

On account of strong affinity of CN^- with cupric ion or gold ion, they can form stable coordination compounds $[\text{Cu}(\text{CN})\text{X}]^{n-}$ and $[\text{Au}(\text{CN})\text{X}]^{n-}$, utilizing the mechanisms to design fluorescent molecular probes are being paid close attention to. The most important advantage of this kind of probe is to realize detection of CN^- in water. Zhai Qiuge[18] selects the coordination compound of Pd^{2+} combined with molecular wire polymeride PDPC which contains pyridine functional group as research system.



Scheme 1 Structures of related fluorescent molecular probes



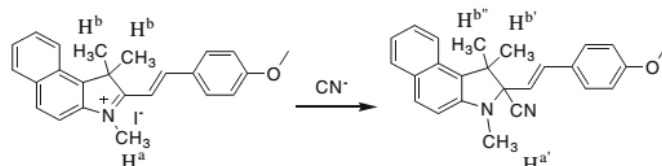
As CN^- has stronger combination ability with Pd^{2+} , PDPC can be released from coordination compounds and recovers the fluorescence of PDPC. According to the variation of fluorescence intensity, the selective detection of CN^- can be done, the lower detection limit is 0.0026mg/L. Huang Xianhong[19] develops a new method to detect CN^- in water, which utilizes the difference of combination ability of CN^- and cupric ion combined with oligonucleotides chain(FAM-DNA) marked with fluorescein. An oligonucleotides chain marked with fluorescein is selected, then it coordinates with cupric ion to form a stable coordination compound (FAM-DNA- Cu^{2+}), thus the fluorescence of the system is canceled. Due to owning stronger combination ability with Cu^{2+} , when CN^- is added into the system, it coordinates with Cu^{2+} , which renders FAM-DNA be released from coordination system, and recovers the fluorescence of system. The detection limit of CN^- in water is 0.00052mg/L, possessing higher selectivity and potential value of application. Li Chao[20] improves the solubility of beta-cyclodextrin in water by covering squaric acids(SQ) with beta-cyclodextrin. In the covered compound, the 2 benzene rings of clathrate compound are implanted into 2 different vacant intracavities, while middle four-membered ring is naked outside. Squaric acids(SQ) is a ideal probe for CN^- because of its intensive and narrow absorbency in near infrared region, as well as its specific sensitivity to cyanides. The coverage of β - beta-cyclodextrin renders insoluble SQ can selectively identify CN^- in pure water system without chemical modification.

After CN^- is added, SQ 634nm is color fading, the logarithm of absorbance of SQ presented good linear relation with concentration of CN^- (the concentration of CN^- from 0.26mg/L to 8.32mg/L).

2.3 Fluorescent molecular probe based on reactions type

With weak alkalinity and strong nucleophilicity, CN^- is utilized to design chemical reaction type fluorescence molecular probe, which has been a hot research subject. In recent years there have been many related reports, that in accordance with the additive subjects, the probes can be classified into addition reaction of C=C bond, C=O bond, C=N bond, as well as lacking of electric center of boron. Lin[21] develops a kind of new CN^- colorimetric and 'turn on' fluorescent probe P3, which is free of disturbance by other anions in THF aqueous solution. The identifying mechanism of CN^- could be explained as: cyanoethylene unit in β site of reagent reacted with CN^- and the additive product of nucleophilic addition owned π bond which are conjugate bond. The minimum detection limit of CN^- is 0.17mg/L, and the method being of low sensitivity, only can be used to monitor sewage discharge. Li[22] develops a fluorescent molecular probe P4, which is based on a colorimetric reaction of 7-trifluoroacetamidocoumarin and CN^- . With simple structure and fast response, P4 can detect CN^- in aqueous solution companied by a few disturbances from other coexisted anion. The ion identifying mechanism is: CN^- reacts with the probe to form 1:1 chemometry formula by nucleophilic impacting of carbonyl group of the probe, the process has been proved by the research of working curve, $^1\text{H-NMR}$ and mass spectrum. The enhanced fluorescence intensity of the probe, supported by DFT/TDDFT calculating, is mainly caused by the improvement of ITC. The detection limit is 0.0078mg/L, response time is less than 30sec., thus, the probe can be used to detect CN^- in water. Yu[23] designs and synthesizes a new ratio fluorescent probe P5, N-2-anthracyltrifluoroacetamide. The anthracyl fragment of P5 works as fluorophore, while trifluoroacetamide works as the acceptor for CN^- . Compared to other reported CN^- probe of trifluoroacetamide derivatives, P5 presents higher sensitivity, and the differences are obvious, CN^- can be identified even if the concentration is as low as 0.1mg/L. Kim[24] utilizes a simple chemometry sensor P6, which uses aldehyded coumarin as Micheal acceptor with double activity to identify CN^- in water. The Micheal additive reaction of P6 and CN^- has presented good selectivity and higher sensitivity. With the CN^- added, the color of the probe changes with the variation of fluorescence, mmol level CN^- can be identified by the naked eyes. Jin Jiayu[25]

synthesizes bi(vinylthiophene) probe P7. In the motivation of UV light, P7 does not emit fluorescence, which means the adding of dicyanoethylene completely cancels the fluorescence of triphenylamine, which may be caused by the inactivation of photoinduction electron transfer (PET). When CN^- is added, CN^- impacts dicyanoethylene, nucleophilic addition is occurred and anions are produced, PET is terminated and the fluorescence of donor can be recovered and enhanced. Thus the fluorescence emission spectrum is same as triphenylamine. The detection limit of CN^- is 0.45 $\mu\text{g/L}$. Sun[26] develops a high selectivity of 'turn-on' type fluorescent probe P8, P8 is based on CN^- being loaded on benzoindoles as fluorescent group and bonding site. CN^- is identified by the nucleophilic reaction of CN^- and benzoindoles kation group, the process is companied with the obviously fluorescence enhanced and color changed. P8 has higher selectivity for CN^- in water, pH range 6-9, detection limit is 0.0012mg/L. The probe can be used to identify CN^- in drinking water, and a set of detection paper is developed.

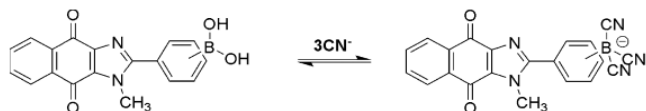


Scheme 2 the mechanism of P8 probe identify CN^-

Mashraqui[27] designs and synthesizes a new CN^- chemometry probe P9, in DMSO, P9 presents selective colorimetric and 'turn-on' fluorescent response to CN^- . When CN^- exists, the red shift is observed in 108nm, as well as 5 times enhancing in fluorescence, which is proved by $^1\text{H-NMR}$. That is an addition reaction process, not only the photoelectron and energy transformation is same, but also the ICT fluorescence enhancing of P9-CN additive product is same. Common anions such as F^- , Cl^- , AcO^- , SCN^- , HSO_4^- , NO_3^- , Br^- , I^- and H_2PO_4^- have no influence or no obvious disturbance in CN^- identifying.

Jamkratoke[28] envelopes a kind of new fluorescent qualitative and quantitative method which utilizes naphthoquinone imidazolyboronic acid fluorescent probe P10(m-NQB and p-NQB), together with cationic surface active agent (cetyltrimethyl ammonium bromide) to identify CN^- in water, and the mechanism shows as scheme. Due to the micro emulsion of CTAB, the boric part of P10 probe gives micro-mol level quantitative detection limit for CN^- , thus the micro emulsion system presents very high selectivity for detection of CN^- . In best conditions, in 460nm, linearity range of fluorescent intensity of m-NQB and p-NQB is

0.013mg/L~39mg/L, detection limit is 0.013mg/L, P10 can be used to identify CN^- in water.



Scheme 3 the mechanism of P10 probe identify CN^-

2.4 The analysis of applicability for fluorescent molecular probe used in environment monitoring

Cyanides belong to second species pollutants in Chinese national integrated wastewater discharge standard, the highest discharge concentration is 0.5mg/L[29]; the level 3 standard of CN^- in surface water is 0.2mg/L[30]; the level 3 standard of CN^- in underground water is 0.05mg/L[31]; the standard limiting value of CN^- in destruction of Japanese abandoned chemical weapons is 0.01mg/L[32]; the limiting value of CN^- in drinking water is 0.05mg/L[33], which is same with drinking water standard recommended by WHO. In China, the recommended detection methods for CN^- are 'isonicotinic acid-pyrazolone method' and 'pyridine-barbituric colorimetric method', the detection limit are respectively 0.004mg/L and 0.002mg/L. Based on the above analysis, only the fluorescent molecular probe being based on reaction type, giving response in aqueous solution can be used to identify the CN^- contained in above-mentioned water types.

3 OUTLOOK

Though the researches of fluorescent molecular probes for cyanide ion are very hot subjects, more of them focus on theoretic research, lacking of practical application. H-bond interaction type probes are weak in sensitivity; coordination type probes have good solubility in water but the sensitivity cannot meet the needs of water quality detection; reaction type probes have no problem in selectivity and sensitivity, but the solubility is poor in aqueous solution, and more of them have to be conducted in organic solution, sometime the reaction are relative slow. Thus developing some fluorescent molecular probes[34] for cyanide ion with characteristics of good selectivity, high sensitivity, good solubility, fast response and practical application are the goals of future research.

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